Nomenclature Rules Summary:

How do I know if it's an acid, ionic compound, or covalent compound?

- If it has a hydrogen on the front, it's probably an acid
- If it has no hydrogen on the front and both elements are to the right of the "staircase": covalent
- If no hydrogen on front and elements are on opp. sides of staircase or there's a polyatomic group: ionic

Ionic Compounds **Formula ! Name** **1.** Write down names of elements (or polyatomic ions) **2.** If first element is d block, write down # (as roman numeral) from charge **3.** Change ending to -ide (unless second half is from polyatomic ion chart)	**Ionic Compounds** **Name ! Formula** **1.** Write down symbol for each element (or polyatomic ion) **2.** Write down charges for each (if d block, use roman numeral for charge) **3.** Use subscripts to balance charges (if polyatomic ion needs subscript, put parentheses around it)
Covalent Compounds **Formula ! Name** **1.** Write down names of elements **2.** Use subscripts to assign number prefixes (if first element is "mono-", then drop the "mono-" **3.** Change the ending to -ide	**Covalent Compounds** **Name ! Formula** **1.** Write symbols for each element **2.** Use number prefixes to assign subscripts
Acids **Formula ! Name** **1.** If binary acid (just H and one other element), then hydro_____ic acid **2.** If oxyacid (group from polyatomic ion chart) and group is "hypo___ite", then hypo___ous acid **3.** If oxyacid and group is "-ite", then ____ous acid **4.** If oxyacid and group is "-ate", then ___ic acid **5.** If oxyacid and group is "per___ate", then per___ic acid	**Acids** **Name ! Formula** **1.** If "hydro____ic acid", then write "**H**" and symbol for element. **2.** If "_____ic acid", then write "**H**" and symbol for "-ate" group from polyatomic ion chart **3.** If "_____ous acid", then write "**H**" and symbol for "-ite" group from polyatomic ion chart **4.** Give **H** subscript to balance charge of element or polyatomic ion group